

Chemistry Chapter 10 States Of Matter

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Chapter 10 - States of Matter Chapter 10 covers an overview of the three most common states of matter: solids, liquids, and gasses. Whereas much of the semester so far has focused on analytical...

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Match description on the right to the correct crystal type on the left
Modern chemistry chapter 10 review answers states of matter section 1.
b ionic crystal (a) has mobile electrons in the crystal
c covalent molecular crystal (b) is hard, brittle, and nonconducting
a metallic crystal (c) typically has the lowest melting point of the four

Chemistry Chapter 10 States Of Matter

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CHEMISTRY NOTES - Chapter 10 States of Matter

Title: Modern Chemistry Chapter 10 States of Matter
1 Modern Chemistry Chapter 10 States of Matter
2 Kinetic-Molecular Theory. The kinetic-molecular theory of matter is based on the idea that particles of matter are always in motion and this motion has consequences that affect its physical properties.
3 Kinetic-Molecular Theory of Gases

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Chapter 10 STATES OF MATTER Liquids Key Terms 1. Fluid 2. Surface tension 3. Capillary action 4. Vaporization 5. Evaporation 6. Freezing Main Ideas 1. The intermolecular forces of liquids determine their properties. The intermolecular forces of liquids determine their properties.

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Modern Chemistry Chapter 10 States of Matter Kinetic-Molecular Theory The kinetic-molecular theory of matter is based on the idea that particles of matter are always in motion and this motion has consequences that affect its physical properties. Kinetic-Molecular Theory of Gases 1- Gases consist of large numbers of tiny particles that are far...

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CHAPTER 10 REVIEW States of Matter SECTION 4 SHORT ANSWER Answer the following questions in the space provided. 1. a When a substance in a closed system undergoes a phase change and the system reaches equilibrium, (a) the two opposing changes occur at equal rates. (b) there are no more phase changes. (c) one phase change predominates.

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View Chapter 10 Section 3_ States of Matter- Chemistry - Google Docs.pdf from BIO 101 at Ganado High School. Chapter 10 Section 3: States of Matter Crystalline solids: A solid that consists of

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Modern Chemistry Chapter 10 States of Matter. Modern Chemistry Chapter 10 States of Matter Kinetic-Molecular Theory The kinetic-molecular theory of matter is based on the idea that particles of matter are always in motion and this motion has

consequences that affect its physical properties. Modern Chemistry Chapter 10 Review Answers States Of Matter

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CHAPTER 10 REVIEW States of Matter SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. Identify whether the descriptions below describe an ideal gas or a real gas. ideal gas a. The gas will not condense because the molecules do not attract each other. ideal gas b. Collisions between molecules are perfectly elastic. real gas

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Modern Chemistry Chapter 10 States of Matter Kinetic-Molecular Theory The kinetic-molecular theory of matter is based on the idea that particles of matter are always in motion and this motion has consequences that affect its physical properties.

Modern Chemistry Chapter 10 Review Answers States Of Matter

Chemistry Chapter 10 States of Matter. STUDY. PLAY. Kinetic-molecular theory. A theory that explains that the behavior of physical systems depends on the combined actions of the molecules constituting the system. Ideal gas. An imaginary gas whose particles are infinitely small and do not interact with each other.

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CHEMISTRY NOTES - Chapter 10 States of Matter Goals : To gain an understanding of : 1.. The kinetic theory of matter. 2. Avogadro's hypothesis. 3. The nature of liquids, solids and gases. 4. Phase changes. NOTES: The kinetic theory of matter states that particles which make up all types of matter are in constant motion.

Modern Chemistry Chapter 10 States Of Matter Test B

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